

Jun 13, 2564 BE - Solution: Let R_e be the radius of the Earth $\hat{=}$ 6377 km. Then the total mass of air in the atmosphere is equal to $2 \text{ t avg avg } e \text{ } 3 \text{ } 2 \text{ m dVol (Air Vol) } 4 \text{ R (Air) ... Read more}$ Jun 13, 2564 BE - Solution: Let R_e be the radius of the Earth $\hat{=}$ 6377 km. Then the total mass of air in the atmosphere is $2 \text{ t avg avg } e \text{ } 3 \text{ } 2 \text{ m dVol (Air Vol) } 4 \text{ R (Air) ...}$ Then the volume of air $V = R \hat{=} \dots 4 \hat{=} R/T = 6377 \hat{=} \dots 4 \hat{=} \hat{=} \dots$
 $3.14 \hat{=} \dots 6.37 = 203.74 \text{ (m}^3\text{)}$ Hide $R_e = 1 \text{ m}^3$ of air. The equation of state for ideal gas for volume $V = 3 \text{ m}^3$ and pressure $p = 1 \text{ atm}$ gives: $R_e = 2 \cdot 10^3 \text{ m}^3$ and $V = 3 \cdot 10^3 \text{ m}^3$. Read more $R_e = 1 \text{ m}^3$ of air. The ideal gas equation of state for volume $V = 3 \text{ m}^3$ and pressure $p = 1 \text{ atm}$ gives: $R_e =$



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